

**UNIVERSITY OF MYSORE
MANASAGANGOTRI, MYSORE 570006**

Course structure and revised syllabus

For

M.Tech Program in Materials Science

2018-2019

Course Structure and Revised Syllabus:

M.Tech. Program in Materials Science - 2013

Admission Eligibility Qualification	Degree after the Program	Hard Core credits (including project)	Soft core credits	Total credits	No. of Years
B.Sc (Physics, chemistry, Geology, Polymer, Biochemistry, Biotechnology, Microbiology, Botany, Zoology, Environmental science, Sericulture, Computer science, Food Science, Mathematics, Statistics, Forensic Science)	M.Tech.	60	46-50	110	3 years
B.Sc(Physics, chemistry, Geology, Polymer, Biochemistry, Biotechnology, Microbiology, Botany, Zoology, Environmental science, Sericulture, Computer science, Food Science, Mathematics, Statistics, Forensic Science)	M.Sc. (optional exit)	50	22-26	76	2 years
B.E/ MBBS/ B.Tech/ B.Pharma/ BDS/ BSc(Ag)(any branch)	M.Tech.	52	12	64	2 years
M.Sc(Physics, chemistry, Geology, Polymer, Biochemistry, Biotechnology, Microbiology, Botany, Zoology, Environmental science, Sericulture, Computer science, Food Science, Mathematics, Statistics, Forensic Science, Nanotechnology)	M.Tech.	52	12	64	2 years

Terms and Conditions:

Total number of seats : 30

Admission is purely based on all India basis.

Admission is purely based on the marks obtained in the entrance exam.

(Syllabus will be specified by the board of studies. The question paper for entrance examination contains four parts each containing Physics, chemistry, Mathematics and Biology. Out of which the candidate can opt for any two parts.)

Other admission regulations are as per the CBCS regulations of University of Mysore 2010.

Project internship can be pursued in any reputed lab or industry or institution in India or abroad.

Odd semester will be during August/September to December/Jan

Even semester will be during Feb/March to June/July. Valuation process is continuous and internal.

This course is essentially an M.Tech Program like any other M.Tech Programs and provision for open electives is not created explicitly. However, if any candidate is interested, he/she can avail the benefit of registration for extra credits to take up open electives.

As and when the new soft core papers are approved by the BOS, the advantage is extended to the students in the higher semester as well along with the incoming students.

Scheme of Evaluation

Each student shall be evaluated continuously by means of tests and/or assignments. There will be two tests: C1 in the sixth week, C2 in the twelfth week and C3 in sixteenth week of the semester. A final examination C4 will be held around eighteenth week.

The C1, C2 and C3 tests will be for 20 marks each and C4 will be for 40 marks. If marks scored in C1+C2+C3 is less than 30%, the student is not allowed to take C4. If the attendance is less than 75% the student is not allowed to take C4.

C1 C2 and C3 tests will be for 1 hour duration. C3 examination will be for 2 hrs duration. Practical tests/examinations will be for 4 hours respectively.

Make up Examination and Re-Appearances for C3 Examination

If a student scores less than 30% in C4, he can take the make-up examination before one month after the announcement of the results with prior information to the department.

If he/she again fails to clear the examination, such students can Re-register for that paper/papers in the department and attend the classes with the regular students of odd and even semesters respectively. This is allowed up to VIII semesters minus the IV semesters of the course (double the duration of the course)

Note: While Re-registering, make sure that per semester with re registration shall not exceed a total of 24 credits for the odd and even semesters.

Hard core courses for M.Tech. Materials Science program.

CODE	Paper title	Core	Credits			
			Lecture	Tutorial	Practical	Total
MSH-1	Introduction to Materials (I-Sem)	Hard	3	1	0	4
MSH-2	Thermodynamics and Statistical Mechanics(I-Sem)	Hard	3	1	0	4
MSH-3	Materials Preparation Techniques (II-Sem)	Hard	3	1	0	4
MSH-4	Methods of Materials Characterization(II-Sem)	Hard	3	1	0	4
MSH-5	Physics and Chemistry of Materials (III-Sem)	Hard	3	1	0	4
MSH-6	Materials and Environmental Effects (III-Sem)	Hard	3	1	0	4
MSH-7	Characterisation lab-1	Hard	0	1	3	4
MSH-8	Characterisation lab-2	Hard	0	1	3	4
MSH-9	Characterisation lab-3	Hard	0	1	3	4
MSH-10	Characterisation lab-4	Hard	0	1	3	4
MSH-11	Characterisation lab-5	Hard	0	1	3	4
MSH-12	Minor Project	Hard	0	1	5	6
MSH-13	Major Project (Only for M.Tech. exit)	Hard	0	2	8	10
			Total : (M.Tech-exit) (For B.E. background)			52
			For B.Sc., background			60

Soft core courses for M.Tech. Materials Science program

Code	Paper title	Core	Credits			
			Lecture	Tutorial	Practical	Total
MSS-1	Structure, Property and Functions of Materials	Soft	3	1	0	4
MSS-2	Basics of Chemistry	Soft	3	1	0	4
MSS-3	Spectroscopic Techniques for Materials	Soft	3	1	0	4
MSS-4	Nanoscale Devices	Soft	3	1	0	4
MSS-5	Nanochemistry	Soft	3	1	0	4
MSS-6	Carbon Nanotubes	Soft	3	1	0	4
MSS-7	Materials for Aerospace Applications	Soft	3	1	0	4
MSS-8	Composite Materials	Soft	3	1	0	4
MSS-9	Polymer Science and Cell Biology	Soft	3	1	0	4
MSS-10	Physics and Chemistry of Polymers	Soft	3	1	0	4
MSS-11	Nano-biotechnology in Health Care	Soft	3	1	0	4
MSS-12	Nano-photonics	Soft	3	1	0	4
MSS-13	Thermodynamic Modelling of Systems	Soft	3	1	0	4
MSS-14	Basics of Engineering Drawing and Graphics	Soft	2	2	0	4
MSS-15	Ceramics Science and Technology	Soft	3	1	0	4
MSS-16	Materials for Renewable Energy and Storage	Soft	3	1	0	4
MSS-17	Basics of Nanotechnology	Soft	3	1	0	4
MSS-18	Enterprise Architecture	Soft	3	1	0	4
MSS-19	Chemical Engineering	Soft	3	1	0	4
MSS-20	Advanced X-ray Diffraction	Soft	3	1	0	4

	Studies					
MSS-21	Analytical and Inorganic Chemistry	Soft	3	1	0	4
MSS-22	Semiconductor Optoelectronics	Soft	3	1	0	4
MSS-23	Analytical and Inorganic Chemistry	Soft	3	1	0	4
		Total soft credits available to choose				66

MSH-1: Introduction to Materials: 4 Credits (Hard core) Total: 48 hrs

Unit-I: Materials through ages

What is Materials Science?, The Relationship of Science and Technology, How is Basic Science Linked to Everyday Materials , A Short History of Materials Science, Classification of Materials, Advanced Materials, Modern Materials' Needs. Materials through ages - Palaeolithic to Modern period. Bulk, fine and Nanomaterials, Crystalline - Functional materials, Smart Materials, Centrosymmetric and Non-centrosymmetric materials, Amorphous, Glasses, Metals, Alloys, Semiconductors - types of semiconductors, Opto-electronic, Polymers, Ceramics, Bio-materials Polymer, Blends, Composites.

12 hours

Unit-II: Crystal Structure

The Structure of Crystalline Solids, Fundamental Concepts of Unit Cells, Metallic, Crystal Structures Density Computations, Polymorphism and Allotropy, Crystal Systems, Point Coordinates, Crystallographic Directions, Crystallographic Planes, Linear and Planar Densities, Close-Packed Crystal Structures, Crystalline and Noncrystalline Materials, Single Crystals, Polycrystalline Materials.

12 hours

Unit-III: Crystal Diffraction and Reciprocal Lattice (Qualitative)

Bragg's Law, Laue Equations, Reciprocal Lattice, Bragg's Condition, Brillouin Zones, Atomic Scattering, Geometrical Structure Factor, Experimental X-Ray Diffraction, Methods of Crystal Structure, Laue Method, Rotary Crystal Method, Powder Method or Debye-Scherrer Method, Weber-Feckner Method.

12 hours

Unit-IV: Chemical Bonding in crystals and Imperfection

Chemical bonding: Ionic bonding, covalent Bonding, Metallic Bonding, Van Der Waals Bonding. Atomic and Ionic radii, Pauling - Ahrens radii of common atom and ions. Spheres in closest packing Cubic closest packing, Hexagonal closest packing, Body centered cubic packing. Voids in closest packing. Crystal defects – point, line, surface /volume defects.

12 hours

References

1. Murugesan R, Modern Physics, S. Chand & Company, 9/e.Rev. Edn. 2003, EN8305SC011.
2. Arthur Beiser, Modern Physics, Addison Wesley Longman Publishing Co (January 1968) ISBN-10: 0201005158
3. Leonid Azaroff, An Introduction to Solids, McGraw-Hill Companies; New edition (1984)
4. Charles Kittel, Introduction to Solid State Physics, John Wiley & Sons, UK 7th Edition (1995)
5. Donald E. Sands, Introduction to Crystallography, Dover Publications, (1994).
6. Darrell Irvine and Nicola Marzari, Fundamentals of Materials Science, MIT Open Course Ware Publications (2005).

MSH-2: Thermodynamics and Phase Equilibria 4 Credits (Hard core) Total: 48 hrs

Unit-I: Kinetic Theory and Gas Laws

Kinetic Theory of Matter, Different States of Matter, Concept of Ideal or Perfect Gas, Kinetic Theory of Gases, Expression for the Pressure of a Gas, Kinetic interpretation of Temperature.

12 hours

Unit-II: Equation of State

Derivation of Gas Equation, Derivation of Gas Laws, Avogadro's Hypothesis, Graham's Law of Diffusion of Gases, Degree of Freedom & Maxwell's Law of Equipartition of Energy, Mean Free Path, Van-der Waals Equation of State, Critical Constants, Corresponding States, Critical Coefficient.

12 hours

Unit-III: Phase equilibrium

Basic concepts of phase, components, Degrees of freedom. Concept of equilibrium and equilibrium constant, Phase rule and its application to One, two and three component systems. Phase diagram and phase equilibrium studies.

12 hours

Unit-IV: Laws of Thermodynamics

Basic concepts and conventions, system and surrounding, Macroscopic physical properties, Time dependent and time independent processes, First law of thermodynamics – expressions for heat and work, Enthalpy, Heat capacity, heats of formations and heats of reactions. Second law of thermodynamics, types of disorders and concept of Entropy and derivation of expressions of entropy. Gibbs free energy, derivation of expression of free energy as function of P and T, Clapyeron equation, partial molar quantities, chemical potential, fugacity and activity. Application of thermodynamics for phase equilibrium studies.

12 hours

References

1. Richard E. Sonntag and Claus Borgnakke, Introduction to Engineering Thermodynamics, Wiley; 2 edition (March 3, 2006), ISBN-10: 0471737593.

2. Ken A. Dill and Sarina Bromberg, Molecular Driving Forces: Statistical Thermodynamics in Chemistry and Biology Garland Science. (Taylor & Francis Group), 2003.
3. Statistical Mechanics by SL Gupta and V Kumar (2011) Pragati Prakashan, India

MSH-3: Materials Preparation Techniques 4 credits (*Hard core*) Total 48 hours

Unit-I

Solid state routes; Nucleation; Role of impurities; Mechanical mixing; Grinding; Solid solution techniques. Top-Down reactions. Rate of crystallization. High temperature processes – heating, annealing, sintering treatment; Sputtering, Spin Coating.

12 hours

Unit -II

Evaporation; precipitation; Solution growth; Nucleation; Rate of crystallization; Supersturation; Top seeded solution growth; sol-gel techniques; high temperature solution; Hydrothermal; Solvothermal methods; Ammonothermal method; Glycothermal; Melt methods- super cooling, Czochorlskii methods; Skull melting.

12 hours

Unit -III

Vapour phase methods - Thin films, epitaxial growth, substrates selection, carrier gases, metastable growth of materials. Chemical Vapour Deposition - Principles, apparatus, examples of CVD growth of thin films, advantages and disadvantages; Chemical Vapour Transportation; Molecular Beam Epitaxy, Liquid Phase Epitaxy, Vapour growth of Nitrides. Metal-organic Vapour phase epitaxy. Plasma Energetics; Laser ablation.

12 hours

Unit -IV

Biological synthesis, Biomimetic method, bacterial synthesis of nanoparticles; Electrochemistry - solvent selection, apparatus, deposition, growth of thin films, coatings, examples; Multi-energy processing - Mechanochemical; Sonochemical; Photochemical; Biochemical, Microbial, Organic synthesis. Growth of organic crystals.

12 hours

References:

1. Springer Handbook of Crystal Growth; Eds: G. Dhanraj, K. Byrappa, V. Prasad, M. Dudley, Springer Verlag (2010)
2. Springer Handbook of Nanotechnology; Eds: Bharath Bhushan, Springer Verlag, 2nd Edition (2009).
3. Handbook of Crystal Growth, K. Byrappa and T. Ohachi, Springer-Verlag (2003).
4. Growth of Single Crystals, R.A. Laudise, Prentice-Hall (1973).
5. Growth and Characterization of Technologically Important Crystals, By: K. Byrappa, R. Fornari, T. Ohachi, H. Klapper, Allied Sciences, New Delhi (2003).

MSH-4: Methods of Materials Characterization 4 credits (Hard co Total: 48 hrs

Unit-I

Thermal analysis; TGA; DTA; DSC; Basic principle. Differences between DTA and DSC. Instrumentation-power compensated DSC, heat flux measurement in DSC. Applications testing the purity and characterizations of various materials (Polymeric, Pharmaceutical, Agricultural materials) Thermo mechanical analysis, Dynamic mechanical analysis, dilatometry; (Thermal expansion) Principles and applications.

12 hours

Unit-II

Microscopic techniques, Phase Contrast Microscopy – Principles, types, hot stage microscopy, Electron imaging techniques; Scanning Electron Microscopy; principle, instrumentation, measurement and analysis; Tunnelling Electron Microscopy; principle, instrumentation, measurement and analysis, Field Emission SEM; Scanning Tunnelling Microscopy; Atomic Force Microscopy; principle, instrumentation, measurement and analysis; Scanning Probe Microscopy; High Resolution TEM; principle, instrumentation, measurement and analysis, High Resolution SEM; principle, instrumentation, measurement and analysis.

12 hours

Unit-III

Particle size measurement; Basic principle of particle analysis, equivalent sphere model, $D[1,0]$, $D[3,2]$ and $D[4,3]$ representations, conversion between length and volume/mass means, mean, median and mode statistics; Methods of measurement of particle size, XRD, optical and laser scattering techniques; surface area and porosity; definition and meaning, measurement using BET method, adsorption isotherms, DC polarization, AC impedance measurements.

12 hours

Unit-IV

Photoluminescence, principle of working, instrumentation and measurement; demonstration of band gap measurement using PL spectrometer; Positron Annihilation Lifetime Spectroscopy; Basics of positron annihilation, the sources of positrons, three methods of positron annihilation techniques; lifetime, Doppler broadening and angular correlation methods, Application of lifetime spectroscopy for the free volume determination in polymers and polymer nanocomposites. Non-linear electro-optical properties of materials, mechanical properties, tensile strength, micro hardness; zeta potential.

12 hours

References

1. Sam Zhang, Lin Li and Ashok Kumar, Materials Characterization Techniques, CRC Press, (2008)
2. Yang Leng, Materials Characterization: Introduction to Microscopic and Spectroscopic Methods, Wiley & Sons (2008)
3. Elton N. Kaufmann, Characterization of Materials, Vol.1, Wiley & Sons (2003)
4. R.A. Laudise, Growth of Single Crystals, Prentice Hall, (1973)

5. G. Dhanaraj, K. Byrappa, V. Prasad and M. Dudley (Eds.), Springer Handbook of Crystal Growth, Springer-Verlag (2010)
6. Peter E.J. Flewitt and R.K. Wild, Physical Methods of Materials Characterization, 2nd Edition, Taylor & Francis (2003)

MSH-5: Physics and Chemistry of Materials 4 credits (Soft core) Total: 48 hrs

Unit-I

Electromagnetic spectrum, accelerator beams, synchrotron radiation, Fermi energy in metals and semiconductors, Specific heat of solids, Conductivity and susceptibility of solids, Energy gap of semiconductors, Hall effect in metals and semiconductors, Quantum theory of metals, Quantized Hall effect.

12 hours

Unit-II

Physical properties of materials Conductivity, density, refractive index, tensile strength, microstructure of bulk, small and nanoscale materials. Optical, Magnetic, Electronic, Semiconducting, Superconducting, Thermal, Electro-optic, Thermo-optic, Superionic, insulator properties of materials.

12 hours

Unit-III

Statistical analysis of Analytical Data, measures of central tendency -Averages and moving averages, Measures of dispersion, Moments, skewness and kurtosis. Frequency distribution and Measures of central tendency using grouped data correlation and regression analysis error estimation and curve fitting. Tests of significance / hypothesis, t-test, f-test and χ^2 test. Probability theory.

12 hours

Unit-IV

Introduction, MX (NaCl, CsCl, ZnS) and MX₂ (fluorite, rutile-cristobalite and cadmium iodide) types. The perovskite and spinel structures. Thermodynamics of ionic crystal formation. Lattice energy, Born-Haber cycle, Born-Landé equation. Applications of lattice energetics. Ionic radii, factors, affecting the ionic radii, radius ratio rules. Electron deficient compounds: Diborane and its reactions, higher boranes, polyhedral boranes (preparations, properties, structure and bonding). Wade's rules, carboranes and metallocarboranes.

References

1. Pillai, Solid State Physics, Narosa Publication, India (2007)
2. J.I. Gersten and F.W. Smith, The Physics and Chemistry of Materials, Wiley & Sons (2001)
3. A. Navrotsky, Physics and Chemistry of Earth Materials, 6th Edition, (1995) Cambridge Series.

MSH-6: Materials and Environmental effects 4 Credit (Hard core) Total: 48 hrs

Unit-I

Basics of Corrosion, Different forms of Corrosion, basis of electrochemical corrosion, theories of corrosion. Local cell theory (Wagner and Traud theory). Current – potential relations (Evan diagram) in corrosion cells. Effect of pH, nature of metal and dissolved oxygen (principle of differential aeration) on corrosion. thermodynamic principles of electrochemical reactions, Electromotive Force Series, Pourbaix Diagrams, Evans Diagrams, Mixed Potential Theory, Passivity,

12 hours

Unit-II

Electrochemical methods to Measure Corrosion: DC Polarization, linear polarization method, AC Impedance; Experimental measurement of corrosion Quantification of corrosion Environmentally Induced Cracking, Corrosion Fatigue, Hydrogen Induced Cracking, Application of Fracture mechanics.

12 hours

Unit-III

Atmospheric Corrosion, Oxidation in Gaseous Environments, Ellingham Diagrams, Role of Protective Scale, Molten Salt Corrosion, Environmental degradation of ceramics, Degradation of Polymeric Materials, Microbial corrosion, Corrosion of Bio-Implants, Corrosion Prevention methods. Corrosion prevention by painting, phosphating and anodic (passivation) and cathodic protection.

12 hours

Unit-IV

Environmental effects from the chemical processes industry (like Pulp mill operations, bleach plants, boilers, paper machine, water treatment plants in the pulp and paper industry and others), infrastructure, and transportation industry.

References

1. D. A. Jones: Principles and Prevention of Corrosion, Macmillan Publ. Co. (1996).
2. C. Scully: The Fundamental of Corrosion, 2nd ed., Pergamon Press: E. E.Stansbury and R. A. Buchanan, Fundamentals of Electrochemical Corrosion, ASM International (2000)
3. M.G. Fontana: Corrosion Engineering, 3rd. Ed., McGraw Hill. (1986)
4. J. M. West: Electrodeposition and Corrosion Control, J. Wiley W. Revie (ed.): Corrosion Handbook, Electrochemical Society Series, John Wiley and Sons (2000).
5. W. Revie (ed.): Corrosion Handbook, Electrochemical Society Series, John Wiley and Sons, 2000: Metals Handbook, Vol. 13: Corrosion, ASM International

MSS-1: Structure, Property and Functions of Materials

4 credits (Soft core)

Total: 48 hrs

Unit-I:Materials

Solid Solutions and Alloys: Phase Transitions: Overview of Crystal Structures: Structure - Property Relations: Neumann's Law: Thermal Properties: Optical Properties: Electrical Properties: Dielectric Properties: Magnetic Properties: Mechanical Properties.

12 hours

Unit-II:Polymers

Importance of polymers. Fundamentals of polymers-Monomers the repeat units, degree of polymerization. Linear, branched and network polymers. Classification of polymers. Polymerization-condensation, addition, free radical, ionic, co-ordination polymerization and ring opening polymerization. Molecular weight and Polydispersion. Average molecular weight concepts-number weight and viscosity average molecular weight. Practical significance of molecular weight. Size of polymer molecules.

12 hours

Unit-III Earths Materials

Basic components of earths Materials, Minerals & rocks- Physical properties of Minerals – density, specific gravity. Silicate structure, Description and classification of minerals – Oxides, Hydroxides Carbonates, Sulphides. and silicates. Solid solutions exsolution, polymorphism and isomorphism, Description of important rock types and ore minerals.

12 hours

Unit-IV:Biomaterials

Protein, antibodies, Enzymes, Nucleotides, immobilization, functionalization methods for biomaterials (Antibody/enzyme/protein), encapsulation methods, biocompatible materials, Biodegradable polymers Bio-compatibility criteria, Toxicity evaluation (Inflammation, hypersensitivity, carcinogenesis). Application of biomaterials in orthopedics (artificial ligaments and tendons), cardiovascular tissues engineering (pacemakers). Biosensors, targeted drug delivery and affordable diagnostics. Carbohydrates, Nucleic acids

12 hours

References

1. Re.E.Newnham, Properties of Materials: Anisotropy, Symmetry, Structure, Oxford University Press.
2. Buddy Ratner. Biomaterials Science. Second edition. Orlando, Academic Press, 2000;
3. Jonathan Black. Biological performance of materials: fundamentals of biocompatibility New York, Marcel Dekker, 1999
4. Joon Park and Joseph Bronzino. Biomaterials: Principles and Applications. Fort Lauderdale FL, CRC Press, 2003.
5. Ferry : Viscoelasticity properties of polymers
6. Aklonies,et al., An introduction of Viscoelasticity in polymers
7. Research and Development on Biosensors for food Analysis in India In: Advances in Biosensors M.S. Thakur and N.G. Karanth, 2003, Oxford University, Press, 2003.
8. From Bioimaging to Biosensors – Noble Metal Nanoparticles in Biodetection, by Lai-kwan Chau (National Chung Cheng University, Taiwan), Huan-Tsung Chang (National Taiwan University, Taiwan)
9. Advances in Biosensors, Volume 1, Reinhard Rennerberg and A.P.F. Turner, 1991.
10. Nanoparticles and Biophotonics as Efficient Tools in Resonance Energy Transfer-Based Biosensing for Monitoring Food Toxins and Pesticides, M. S. Thakur, Rajeev Ranjan, Aaydha C. Vinayak, Kunhitlu S. Abhijith, and Richa Sharma. Advances in Applied Nanotechnology for Agriculture, Chapter 4, 2013, pp 55-84, ACS Symposium Series, Volume 1143.

MSS-2 : Basics of Chemistry

Unit I:

Elements of quantum mechanics: Wave mechanical concept of the atom, dual nature of electron, derivation of de-Broglie's equation. Heisenberg's uncertainty principle and its significance. Schrodinger wave equation-explanation of the terms therein (no derivation) Eigen values and functions, significance of ψ and ψ^2 . Quantum numbers and their significance. Shapes of s, p and d orbitals. Effective nuclear charge, screening effect-based on Slater's rules (problems to be worked out). General energy level diagram of multi electron atom (up to $n=4$). Pauli's exclusion principle, Hund's rule, $(n+1)$ rule, Aufbau principle.

Periodic Table and Periodicity: Classification of elements into s, p, d, and f-blocks, cause of periodicity. Detailed discussion of the following periodic properties of elements with examples: Atomic radius, Ionization enthalpy, Electron gain enthalpy and Electronegativity.

Chemical bonding: Ionic bond, Covalent bond, Coordinate bond, metallic bonding and hydrogen bonding. **12 hours**

Unit II:

Hybridization: Tetravalency of carbon, sp^3 , sp^2 and sp – hybridization (in brief). Bond length, bond angle, bond energy, localized and delocalized chemical bonds – resonance and hyperconjugation effects.

Types of organic reactions: Definition with examples of addition, substitution, elimination, isomerisation, oxidation, reduction, condensation and rearrangement reactions.

Stereochemistry: Introduction, definition, elements of symmetry (plane, centre, simple axes and alternative axes), asymmetry and dissymmetry, Chirality, designation of configuration – R-S notation. Optical activity – explanation – cause of optical activity (non-super impossibility). Enantiomers and diastereomers optical isomerism in tartaric acid and biphenyls, racemisation, resolution, methods of resolution (Chemical and biochemical methods) Walden inversion, asymmetric synthesis (partial and absolute).

Geometrical isomerism: Definition with example, designation of cis-trans and E-Z notations with examples. Geometrical isomerization of aldoximes and ketoximes, **12 hours**

Unit III

Colligative Properties: Concept of vapour pressure, variation of vapour pressure with temperature. Definition of boiling point and freezing point, effect of dissolution of solute on the vapour pressure of the solvent. Lowering of vapour pressure. Raoult's law – relation between relative lowering of vapour pressure and molar mass (to be derived). Determination of relative molar mass of solute by dynamic method. Elevation of boiling point and its relation to lowering of vapour pressure and molar mass (to be derived). Ebullioscopic constant of the solvent and its relation to the boiling point (only equation). Determination of molar mass of the solute by Walker-Lumsden method. Depression in freezing point and its relation to lowering of vapour pressure and molar mass (to be derived). Cryoscopic constant and its relation to the melting point (equation). Determination of molar mass of a non-volatile solute by Beckmann's method (problems to be worked out).

Semi permeable membrane – natural and artificial, preparation of copper ferrocyanide membrane by Morse-Frazer method. Definition of osmosis, osmotic pressure (mention application), determination of osmotic pressure by Berkley-Hartley's method, laws of osmotic pressure analogy with gas laws, determination of molar mass from osmotic pressure measurements (relation to be derived), isotonic solutions, plasmolysis. **12 hours**

Unit IV

Indicator – Definitions, types (acid-base, redox, adsorption indicators), examples for each type. Theory of indicators – Oswald's theory and Quinonoid theory – indicator constant – action of phenolphthalein and methyl orange in acid-base solutions – pH titration curves for strong acid vs strong base, weak acid vs strong base, weak base vs strong acid, choice of indicators in these types of titrations – color change and pH range. Universal indicator – definition.

Chromatography: i. Paper: introduction to ascending, descending and circular, R_f value and its applications; ii. **TLC:** Introduction and applications; iii. **Column Chromatography:** Introduction, principle and experimental details and applications; iv. **Gas Chromatography:** Introduction, apparatus, programmed temperature gas chromatography, quantitative analysis of GLC and v. **HPLC:** Introduction, schematic diagram of instrumentation and application.

12 hours

Reference:

1. Basic Inorganic Chemistry – 3rd edition. F.A. Cotton, G. Wilkinson and P.L. Gaus, John Wiley and Sons (2002).
2. Inorganic Chemistry, 3rd edition. James E. Huheey, Harper and Row Publishers (1983).
3. Inorganic Chemistry, 3rd edition. G.L. Miessler and D.A. Tarr, Pearson Education (2004).
4. Inorganic Chemistry, 2nd edition. D.F. Shriver, P.W. Atkins and C.H. Langford, Oxford University Press (1994).
5. H. Pine, Hendrickson, Cram and Hammond, Organic Chemistry, Mc Graw Hill, New York, 1987.
6. I.L. Finar, Organic Chemistry, ELBS Longmann, Vol. I & II, 1984.
7. R.K. Bansal, Organic Reaction Mechanism, Wiley Eastern Limited, New Delhi, 1993.
8. J. March, Advanced Organic Chemistry, Wiley Interscience, 1994.
9. Physical Chemistry by P.W. Atkins, ELBS, 5th edition, Oxford University Press (1995).
10. Text Book of Physical Chemistry by Samuel Glasstone, MacMillan Indian Ltd., 2nd edition (1974).
11. Elements of Physical Chemistry by Lewis and Glasstone.
12. Fundamentals of physical chemistry – Maron and Lando (Collier Macmillan) 1974.

MSS-3: Spectroscopic Techniques for Materials 4 credits (Soft core) Total: 48 hours

Unit I

UV-Visible spectroscopy: Modes of electronic excitations, simple chromophoric groups—systems of extended conjugation, aromatic systems. Types of auxochromes—functions of auxochromes, absorption and intensity shift. Types of transition probability, types of absorption bands, solvent effects and choice of solvent. Effect of polarity on various type of bonds, Woodward's empirical rules for predicting the wavelength of maximum absorption: - Olefins, conjugated dienes, cyclic trienes and polyenes, α,β -unsaturated aldehydes and ketones, benzene and substituted benzene rings. **12 hours**

IR spectroscopy: Principles, Hook's law, characteristic group frequencies and skeletal frequencies. Finger print region. Identification of functional groups: Alkenes, alkynes, aromatics, carbonyl compounds (aldehydes and ketones, esters and lactones), halogen compounds, sulphur and phosphorous compounds, amides, lactams, amino acids, and imines. Factors affecting group frequencies and band shapes, conjugation, resonance and inductance, hydrogen bonding and ring strain, tautomerism, cis-trans isomerism. Applications of IR spectra to co-ordination compounds, organotransition metal complexes (N,N-dimethyl acetamides, urea, DMSO, NO_3^- , SO_4^{2-} , NO_2^-) **12 hours**

Unit II

Nuclear magnetic resonance spectroscopy: General introduction and definition, magnetic properties of nuclei (magnetic moment, g factor) and theory of nuclear resonance. Larmor precession frequency, resonance condition and relaxation processes.

Chemical shift: Standards employed in NMR, factors affecting chemical shift, electronegativity, shielding and deshielding mechanism, Vander waals deshielding, H-bonding, diamagnetic and paramagnetic anisotropics. Spin-spin coupling, chemical shift values and correlation for protons bonded to carbon and other nuclei. Instrumentation and sample handling.

Equivalence and magnetic equivalence proton exchange reactions, effects of chiral center, complex spin-spin interaction, stereochemistry, hindered rotation, Karplus curve—variation of coupling constants with dihedral angles. Simplification of complex spectra: isotopic substitution, increasing magnetic field strength, double resonance, spin decoupling, constant shift reagents, solvent effect, Fourier-transfer technique, variable temperature profile, nuclear overhauser effect (NOE).

12 hours

Unit III

Mass spectrometry: Principles, instrumentation, different methods of ionization, EI, CI, FD and FAB, ion separators: single focusing separator with magnetic diffraction, focusing analyzer, time-of-flight separator and quadrupole analyzer. Mass spectra: molecular ion, base peak, meta-stable peak, nitrogen rule and McLafferty rearrangement. Mass spectral fragmentation of organic compounds and common functional groups: normal and branched alkanes, alkenes, cycloalkanes, benzene and its derivatives, alcohols, phenols, aldehydes and ketones, carboxylic acids, and their derivatives, amines, nitrocompounds. Determination of molecular formula by accurate molecular weight and isotopic abundance methods. Examples of mass spectral fragmentation of organic compounds with respect to their structure determination.

LC-MS, LC-MS/MS, GC-MS: Principles and applications

Composite problems involving the applications of UV, IR, and ^1H -NMR and mass spectroscopic techniques. **12 hours**

UNIT – IV

Electron Spin Resonance Spectroscopy: Basic principles, hyperfine couplings, the 'g' values, factors affecting 'g' values, isotropic and anisotropic hyperfine coupling constants, Zero Field splitting and Kramer's degeneracy. Measurement techniques and Applications to simple inorganic and organic free radicals and to inorganic complexes.

Mössbauer spectroscopy: The Mössbauer effect, chemical isomer shifts, quadrupole interactions, measurement techniques and spectrum display, application to the study of Fe²⁺ and Fe³⁺ compounds, Sn²⁺ and Sn⁴⁺ compounds(nature of M-L bond, coordination number and structure), detection of oxidation states and inequivalent Mössbauer atoms.

Photoelectron spectroscopy (PES): Brief theory, principle of working, instrumentation and data acquisition analysis. **12 hours**

REFERENCES:

1. Organic Spectroscopy-3rd Ed.-W.Kemp(Paggrave Publishers, New York), 1991.
2. Spectrometric Identification of Organic Compounds - Silverstein,Bassler & Monnill (Wiley) 1981.
3. Spectroscopy of Organic Compounds-3rd Ed.-P.S.Kalsi (New Age, New Delhi) 2000.
4. E.A.V.Ebsworth, D.W.H.Ranklin and S.Cradock: Structural Methods in Inorganic Chemistry, Blackwell Scientific, 1991.
5. J. A. Iggo: NMR Spectroscopy in Inorganic Chemistry, Oxford University Press, 1999. 6. C. N. R .Rao and J. R. Ferraro: Spectroscopy in Inorganic Chemistry, Vol I & II (Academic) 1970
7. Spectroscopy, B. P. Straughan and S. Salker, John Wiley and Sons Inc., New Yourk, Vol.2, 1976.
8. Application of Absorption Spectroscopy of Organic Compounds, John R. Dyer, Prentice/Hall of India Private Limited, New Delhi, 1974.
9. Organic Spectroscopy, V. R. Dani, Tata McGraw-Hall Publishing Company Limited, New Delhi. 1995.
10. Interpretation of Carbon-13 NMR Spectra, F.W. Wehrli and T. Wirthin, Heyden, London, 1976.
11. NMR spectroscopy-Powai

MSS-4: Nanoscale Devices

4 credits (Soft core)

Total: 48 hrs

Unit - I Processing

Silicon Processing methods- Cleaning / Etching- Oxidation- Gettering- Doping-Epitaxy- Sputtering-Chemical Vapor Deposition (CVD), Plasm Enhanced CVD- Reactive Ion Etching (RIE)- Moore's law-Design rules for 45nm, 32nm, and beyond- Semiconductor device roadmap- Silicon -insultaor technology- Gate of high -K dielectrics

12 hours

Unit -II Fabrication

Thermal manufacturing - Rapid Thermal Processing and beyond: Applications in

Semiconductor Processing of Complimentary Metal Oxide Semiconductor. Memory devices – Volatile and Non-volatile memory. The material challenge of Ultra thin body (UTB) - metal–oxide–semiconductor field-effect transistor (MOSFET, MOS-FET, or MOS FET) for technology. Insulated-gate field-effect transistor (IGFET). 3D inter-connect technology.

12 hours

Unit -III Lithographic techniques

Top down approach to nanolithography-Immersion lithography- Optical lithography, UV photolithography- Phase lithography- Including Plasma X-ray sources- E-Beam Lithography- Focused Ion beams- Photoresist. Soft lithography for nanofilms and nanoscale patterning. Lithographic techniques and surface chemistry for the fabrication of PEG-passivated protein microarrays.

12 hours

Unit -IV Fabrication of nanostructures

Si-Ge-Si-C , Diamond,-Synthesis- Defects and properties on the nanoscale-Bottom up approach- Chemical self assembly- Properties of nanoparticles, Nanoclusters, Nanotubes, Nanowires, Nanoflowers, and Nanodots. Invention jet prints nanostructures with self-assembling material. Surfaces, Interfaces and Thin films. Fabrication of nanostructures using polymer brushes. Top down fabrication of nanostructures. Template assisted nanostructure growth. Nanostructured Films on Silicon Surfaces, Deposition, Characterization and Application of Nanocrystalline Diamond Films, Boron Nitride films.

12 hours

References:

- 1) Mark J Jackson, *Micro and Nanomanufacturing* , Springer; First Edition, (2006) ISBN-10:038725744
- 2) Dieter K, Schroder, *Semiconductor Material and Device Characterization*, Wiley-IEEE Press, 3rd Edition, (2006) ISBN-10:0471739065
- 3) L. B. Freund and S Suresh, *Thin film materials: Stress, Defect formation and surface Evolution*, Cambridge University Press, (2004) ISBN-10:0521822815
- 4) Zheng Cui, *Micro-nanofabrication: Technologies and Applications*, Springer First Edition (2006), ISBN-10:3540289224
- 5) R. Kassing, P. Petkov, W. Kulish, C. Popov., *Functional Properties of Nanostructured Materials*. Springer (ISBN: 978-1-4020-4595-0 (Print) 978-1-4020-4594-3 (Online)

MSS-5: Nanochemistry

4 Credits (Soft core)

Total: 48 hrs

Unit-I Fundamentals of molecular self-assembly

The nanoscale and colloidal systems-Fundamentals of Surface and interfacial chemistry- Surface tension and Wettability-Insoluble monolayers-Surface Chemistry and monolayers- Electrostatic interactions in self assembling systems-Self-Assemble of amphiphiles- Monolayers-Micelles and microemulsions-the Structure and properties of Micelles.

12 hours

Unit-II Nanomaterials

Defining nanodimensional materials-Size effects in Nanomaterials-Application and technology development-Supramolecular machines-Fundamentals of energy transfer and photon motion manipulation-Solar energy harvesting-Fundamentals of electron motion manipulation-Electron pumping and molecular wires-General methods available for the synthesis of nanomaterials-Manipulation of Nanoparticles-Nanofabrication-Methods- Bottom up methods- Photolithography-Scanning probe methods-Soft lithography

12 hours

Unit-III Dendrimers and their applications

PPI and PAMAM dendrimers, Synthesis-Generation 1 and Generation 2, Application in Drug targeting

12 hours

Unit-IV Functionalization and applications of Nanomaterials

Chemical functionalization - Recent advances in Thiol-Au and Silane Chemistry- Layer-by-Layer synthesis of multilayer assemblies Applications - Quantum dots - nanocores and applications. Detailed description of the fabrication of functionalised Gold Nanocores and their application in cancer therapy.

12 hours

References

1. The chemistry of nanomaterials Volume 1, Synthesis, Properties and Applications: Edited by CNR Rao, A Muller, A K Cheetham; (2005) John-Wiley and Sons, Inc, ISBN:3-527- 30686-2
2. The chemistry of nanomaterials Volume 2, Synthesis, Properties and Applications: Edited by CNR Rao, A Muller, A K Cheetham; (2005) John-Wiley and Sons, Inc, ISBN:3-527- 30686-2

MSS-6: Carbon Nanotubes

4 Credits (Soft core)

Total: 48 hrs

Unit-I Introduction and Morphology

Carbon allotropes, Fullerenes, Discovery of carbon nanotubes, Single and double walled carbon nanotubes, Bundled Nanotubes. The structure of carbon nanotubes: From a Graphene Sheet to a Nanotube - Achiral and Chiral Nanotubes, theoretical discussion, experimental studies, ZigZag and Armchair Nanotubes - Eulers Theorem in Cylindrical and Defective Nanotubes.

12 Hours

Unit-II Production Techniques of Nanotubes

Carbon nanotube production: Arc discharge, Laser ablation, Chemical vapour deposition, Hydrothermal, Miscellaneous synthesis methods. Growth mechanism of Multi/Singlewalled nanotubes by arc discharge and Chemical vapour deposition method. Role of Metal Catalyst, High- Pressure CO Conversion (HPCO), Experimental Puzzles of MWNT - Aspect Ratio-Perfection-chemical inertness. Purification and processing of MWCNTs and SWCNTs, separation of metallic and semiconducting SWCNTs.

12 Hours

Unit- III Electrochemical properties and functionalization of Nanotubes

Electronic properties of SWCNTs, Electrode potentials versus work functions,

electrochemistry of Carbon nanotubes, cyclic voltammetric investigations of SWCNTs, Standard redox potentials of SWCNTs. Physical properties: theoretical and experimental predictions of mechanical properties, optical properties of nanotubes, raman spectroscopy. Covalent functionalization: Chemical functionalization, Defect group functionalization, Direct sidewall functionalization. Noncovalent functionalization: with small aromatic molecules, heterocyclic polyaromatic systems, polymers, surfactants and ionic liquids.

12 Hours

Unit-IV Applications of Nanotubes

Catalytic: Macroscopic shaping of CNTs, specific metal support interaction, dispersion of the active phase, electrically and thermally conductive supports. Biomedicine: cell penetration, drug delivery, gene delivery, anticancer approaches, antioxidants properties, toxicity. Probes and Sensors: nanotube tips for atomic force microscopy, gas sensors, biosensors, physical sensors. Containers: mechanisms of nanotube filling, fullerene C₆₀, higher fullerenes, endohedral fullerenes, functionalized fullerenes, molecules without metal atoms, organometallic and coordination compounds, oxides and hydroxides, inorganic materials. Effect of Doping on conductivity - Harnessing Field Enhancement - Flat Panel Displays

12 hours

References

1. Dirk M. Guldi and Nazario Martin, Carbon nanotubes and related structures
2. Peter J. F. Harris, Carbon nanotube science.
3. M.Endo, S.Iijima, M.S. Dresselhaus, Carbon Nanotubes, Pergamon; Ist Ed Edition (December 1st 1996), ISBN-10: 0080426824
4. Ado Jorio, Mildred S. Dresselhaus, and Gene Dresselhaus Carbon Nanotubes: Advanced Topics in the Synthesis, Structure, Properties and Applications, Springer; 1 edition (April 20,2001) ISBN-10: 3540410864

MSS-7: Materials for Aerospace Applications 4 credits (soft core) 48 hours

Unit I: Introduction to aerospace materials

The importance of aerospace materials, understanding aerospace materials, the main types of polymer based aerospace materials, advantages of polymer based materials, High performance like light weight and other important properties of materials for aerospace applications including modified structures on nano-scale materials

12 hours

Unit II: Polymers for aerospace structures

Aerospace applications of polymers, advantages and disadvantages of polymers for aerospace applications, Polymerization routes, Thermoset, Thermoplastics and elastomers, Structural adhesives, Mechanical properties, Polymer additives, Polymers for radar absorbing materials, Fracture toughness properties, Ductile/brittle fracture transition for polymers, Improving the fracture toughness of aerospace materials

12 hours

Unit III: Manufacturing of fibre-polymer composite materials

Fibre reinforcements for composites, Production of prepregs and fabrics ,Core materials for sandwich composites, Composites manufacturing using prepregs, Composites manufacturing by resin infusion, Machining of composites , Damage mechanics of composites. Fracture

processes, Stress concentration effects in materials, Fracture mechanics, Application of fracture mechanics to aerospace materials

12 hours

Unit IV: Fibre-polymer composites for aerospace structures and engines

Types of composite materials, Aerospace applications of fibre-polymer composites, Advantages and problems with using fibre-polymer composites, Mechanics of continuous fibre composites, Sandwich composites, Environmental durability of composites Fatigue stress, Fatigue life (S-N) curves, Fatigue crack growth curves, Fatigue of fibre-polymer composites, Creep behaviour of materials, Creep of polymers and polymer composites, Designing Creep resistant materials.

12 hours

References:

1. Mechanics of Fibrous Composites, C.T. Herakovich, John Wiley & Sons, Inc., New York, 1998
2. Analysis and Performance of Fibre Composites, B.D. Agarwal and L.J. Broutman, John Wiley & Sons, Inc. New York.
3. Fundamentals of Modern Manufacturing, Materials, Processing, and Systems, 2nd edition, Mikell P. Grover, John Wiley & Sons, inc
4. Structure and properties of engineering materials, fifth edition, Henkel and Pense, McGraw Hill, 2002
5. Mechanical Testing of Advanced Fibre Composites, J.M. Hodgkinson, Woodhead Publishing Limited, Cambridge, 2000.
6. ASTM standards.

MSS-8: Composite Materials

4 Credits (Soft core)

Total: 48 hrs

Unit-I: Introduction: Definition, Reason for composites, classification of composites, Raw materials, classification, Chemistry, Properties and applications. Matrix: Thermoplastics-Raw materials, Physical and chemical properties, Thermal behaviour and mechanical properties, Thermosets-Epoxy; Curing reactions, Hardener, Gel time Viscosity Modifications, Prepreg making, Unsaturated polyester resin; catalyst, curing reaction, Viscosity modifier, Alkyd Resin, Vinly ester, polyimides, Physical and chemical properties, Thermal behaviour, Mechanical Properties and uses, Elastomeric composites. **12 hours**

Unit-II: Reinforcements; Types, Properties, Uses of silica, Titanium dioxide, Talc, Mica, etc., Flake, Fibres -Structure, property and applications of natural and synthetic fibres, organic and inorganic fibres. Example: Glass, Carbon, Aramid, Nylon, Boron, Aluminium carbide, Silk, Jute, Sisal, Cotton, etc, Coupling agents. **12 hours**

Unit-III: Processing : Thermoplastic, Thermosets, etc., Types of methods, Processing conditions advantages and disadvantages, Film forming, Lamination, Sandwich, etc., Hand Layup methods, compression and Transfer molding, Pressure and Vacuum bag process, Filament winding, Spin coating, Pultrusion, Reinforced RIM, Injection molding of Thermosets, SMC and DMC, Factors affecting the performance of Composites. Testing of Composites; Destructive and non-destructive tests, Tensile, Compression, Flexural, ILSS, Impact strength, HDT, Basic Concepts of fracture mechanism. **12 hours**

Unit-IV: Composite product design, Fundamentals, Definitions, Structure -Material -Design relationships, Design methodologies, Material Considerations, Application of Composites- Aerospace, Transport, marine, Structural, Chemical and Corrosion resistant products, sports, electrical, Electronic, Communication, Biomedical Applications, Repairs and maintenance, etc., Nanocomposites: -Types, preparation, characterization and applications.

12 hours

References

1. Handbook of Composites by G. Lubin, Van Nostrand, New York, 1982.
2. Polymers and Polymer Composites in Construction L.C. Holleway, 1990
3. Engineering plastics and Composites by John C. Bittence, 1990
4. Handbook of Plastics, Elastomers and Composites by Charles A Harper, 1975
5. Designing with Reinforced Composites - Technology Performance, Economics Rosato, 2nd ED. 1997
6. Delwane Composite design Encyclopedia (Vol 3 Processing and Fabrication / Technology_ Ed. Leif A. Carlsson. and Joahn W. Hillispie, Technomic Publishing Ah. Lancaster U.S.A.
7. Fibre Glass Reinforced Plastics Nicholas P. Cheremisinoff and Composites Paul N. Cheremisinoff., Noyes publications, N.J. U.S.A (1995)
8. Composite applications The Future is now, Thomas J. Drozdr, (Eds), Published by Society of Manufacturing Engineers, Michigan, 1989.
9. Polymer Layered Silicate and silica nano Composites, Y.C. Ke, P. Stroeve and F.s. Wang, Elsevier, 2005
10. Hand Book of Plastics Testing Technology Vishu Shah, John Wiley & Sons, Inc NY. (1998)

MSS-9: Polymer Science and Cell Biology

4 Credits (Soft core) Total: 48 hrs

Unit-I: Basic Polymer Science

Polymer synthesis and structure-chain structure and configuration - Amorphous polymer state-Conformation of polymer chain - Macromolecular dynamics - structure of crystalline polymers-polymers in the liquid crystalline state - Glass, Rubber - Transition behaviour

12 hours

Unit-II: Properties of polymers

Methods of measuring transitions in polymers - Cross linked polymers and rubber elasticity - Polymer Visco elasticity and Rheology, Mechanical Behaviour of polymers - Polymer surfaces and interface.

12 hours

Unit-III: Cell Structure

Overview of cells- Origin and overview of cells- Cells as experimental model - Chemistry of cells -Molecular composition- Enzymes as biological catalysis- Metabolic energy- Biosynthesis of cell constituents- Cell membranes - Fundamentals of molecular biology- Hereditary genes and DNA-Expression of Genetic DNA -Recombinant DNA- Detection of nucleic acid and proteins.

12 hours

Unit-IV: Central Dogma of life

Flow of genetic information-Organisation and sequence of cellular genomes, Replicating- Maintenance of genomic DNA- RNA synthesis and processing- Protein synthesis and regulation. Cell structure and function- the nucleus- Protein sorting and transport,

Bioenergetics and metabolism-Cytoskeleton and cell movement- the cell structure -cell regulation-Cell signalling -Cell Cycle-Cancerous cells. **12 hours**

References

1. Leslie Howar Sperling, Introduction of physical polymer science, Wiley-Interscience; 4th Edition (2005) ISBN-10: 047170606X.
2. Geoffrey M Cooper, The cell- A molecular approach, Sinauer Associates Inc; 4 the Edition, (2006) ISBN-10: 0878932194

MSS-10: Physics and Chemistry of Polymers 4 Credits (Soft core) Total: 48 hrs

Unit- I Basic concepts: Classification of polymers, Nomenclature of Polymers concepts such as monomer,polymer,oligomer,dendrimer ; functionality and physical state (amorphous and crystalline) Stereo- regular polymers, co- polymers, block and graft co-polymers, molecular forces and chemical bonding in polymers. Polymerization mechanism, addition and condensation including co-ordination, cationic, anionic, Ring opening Redox polymerization, Living Radical polymerization-Atom transfer radical polymerization.

12 hours

Unit-II Methods of Polymerization – Bulk, solution, precipitation polymerization, Suspensions,emulsion, melt polycondensation, interfacial polymerization, solution . polycondensation, solid phase, gas phase and (formulation, mechanism, properties of the polymer produced advantages and disadvantages of each technique).Criteria of polymer solubility, solubility parameter, thermodynamics and phase equilibria of polymer solution, Fractionation of polymers by solubility.

12 hours

Unit - III Structure property relationship in polymers, configuration of polymer chains, crystallinity in polymers, crystallisation and melting, strain induced morphology, crystalline melting point, glass transition temperature, factors influencing Tg and Tm. Polymer rheology: Viscous flow, Kinetic theory of rubber elasticity, Visco-elasticity.

12 hours

Unit – IV Molecular weight and size: Importance of molecular weight, Molecular weight distribution, concept of Average molecular weights, measurement of molecular weight by end group analysis, colligative property, Light scattering, Ultracentrifugation, dilute solution viscosity, Gel permeating chromatography.

12 hours

References:

- 1.Text book of Polymer Science – Fred W. Billmeyer, J.R. John Wiley & Son, New York.(1990)
- 2.Polymer Science – V.R. Gowarikar, N.V. Vishwanathan, Jayadev Shreedhar Wiley Eastern Ltd. New Dehli, India(1986)
- 3.Analysis of polymers – an introduction – T.R. Crompton. Smithers Rapra Technology Pvt Ltd., SY4 4NR, UK, 2008
- 4.Experimental methods in polymer chemistry – J.F. Rabek. John Wiley and Sons NY(1980)
- 5.Polymer Science, P.L. Nayak, Kalyani publishers, New Dehli.(2005)
- 6.Analysis and Characerisation of Polymers – Sukumar Maiti, Ansandhan Prakashan, Midnapur, India.(1978)

MSS-11: Nano-biotechnology in Health Care 4 Credits (Soft core) Total: 48 hrs
Unit-I

Introduction on biocompatible nano-particles (Gold, silver, zinc, carbon, graphene and quantum dots), synthesis and applications. Nano-biosensors and Lab-on-chip devices for ultrasensitive diagnostics: Amperometric, potentiometric and potentiometric and fluometric biosensors (glucose biosensors, enzyme biosensors, screen printed electrodes). Affordables diagnostics. **12 hours**

Unit-II

Techniques involved in Nanoparticles application in diagnostics and characterization: Fluorescence resonance energy transfer (FRET), Surface energy transfer (SET), Raman light scattering, Surface Plasmon Resonance (SPR), Transmission electron microscopy, Scanning electron microscopy Atomic force microscopy, confocal microscopy, scanning tunneling microscopy, fluorescence spectroscopy. **12 hours**

Unit-III

Micro-array (DNA and Protein array)-concepts and advantages of Microfluidic devices, Materials for manufacture of microfluidic devices, (Silicon and PDMS). Nanoparticles for Optical Imaging of Cancer, Nanogold in Cancer Therapy and Diagnosis, Nanotubes, Nanowires, Nanocantilevers and Nanorods in Cancer Treatment and Diagnosis. Carbon Nanotubes in Cancer Therapy and Diagnosis. **12 hours**

Unit-IV

Introduction, Liposomes, dendrimers, Layer by layer deposition, self-assembled monolayers, In-vivo imaging. Nanobiotechnology for food health and environment. **12 hours**

References:

1. From Bioimaging to Biosensors – Noble Metal Nanoparticles in Biodetection, by Lai-kwan Chau (National Chung Cheng University, Taiwan), Huan-Tsung Chang (National Taiwan University, Taiwan)
2. Advances in Biosensors, Volume 1, Volume 1, Reinhard Renneberg and A.P.F. Tuner, 1991.
3. Nanoparticles and Biophotonics as Efficient Tools in Resonance Energy Transfer-Based Biosensing for Monitoring Food Toxins and Pesticides, M. S. Thakur, Rajeev Ranjan, Aaydha C. Vinayaka, Kunhitlu S. Abhijith, and Richa Sharma. Advances in Applied Nanotechnology for Agriculture, Chapter 4, 2013, pp 55-84, ACS Symposium Series, Volume 1143.
4. Nanomaterials for Cancer Diagnosis, Challa S. S. R. Kumar (Editor), ISBN: 978-3-527-31387-7, 448 pages, January 2007.
5. Microarray Technology and Its Applications, Uwe R. Muller, Dan V. Nicolau, Springer, 30-Mar-2006- Technology and Engineering.
6. Nanoparticulates as Drug Carriers edited by Vladimir P Torchilin (Northeastern University, USA), 2006, Imperial college press.

MSS-12: Nano-photonics**4 Credits (Soft core)****Total: 48 hrs*****Unit-I: Quantum confined materials***

Quantum dots - Optical transitions absorption - interband transitions - quantum confinement - intraband transitions fluorescence/luminescence - photoluminescence/fluorescence optically excited emission electroluminescence emission.

12 hours***Unit-II: Plasmonics***

Internal reflection and evanescent waves - plasmons and surface plasmon resonance - Attenuated Total reflection - Grating SPR coupling - Optical waveguide SPR coupling - SPR dependencies and materials - plasmonics and nanoparticles.

12 hours***Unit-III: New Approaches in Nanophotonics***

Near field optics - Aperture less near field optics - near field scanning optical microscopy (NSOM or SNOM) - SNOM based detection of plasmonic energy transport - SNOM based visualization of waveguide structures - SNOM in nanolithography - SNOM based optical data storage and recovery.

12 hours***Unit-IV: Biophotonics***

Interaction of light with cells-tissues - nonlinear optical processes with intense laser beams - photo induced effects in biological systems - generation of optical forces - optical trapping and manipulation of single molecules and cells in optical confinement - laser trapping and dissection for biological systems-single molecules biophysics - DNA protein interaction.

12 hours**References**

1. H Masuhara, S Kawata and F Tokunga, Nanobiophotonics, Elsevier Science 2007.
2. BEA Sale and A C Teich, Fundamentals of photonics, John Wiley and Sons, New York 1993.
3. M Ohtsu, K Kobayashi, T Kawazoe and T Yatsui, Principles of Nanophotonics (Optics and Optoelectronics), University of Tokyo, Japan (2003).
4. P N Prasad, Introduction to Biophotonics, John Wiley and Sons (2003)

MSS-13: Thermodynamic Modelling**4 Credits (Soft core)****Total: 48 hrs*****Unit-I***

Entropy as a measure of unavailable energy. Entropy change during spontaneous process. Helmholtz and Gibbs free energies. Thermodynamic criteria of equilibrium and spontaneity. Variation of free energy with temperature and pressure. Maxwell's relations, Van't Hoff's reaction isotherm and isochore, Gibbs-Helmholtz equation. Determination of free energy changes. Nernst heat theorem and Third law of thermodynamics-calculation of absolute entropies. Concepts of thermodynamics. Partial molar volume and its determination by density measurements. Order of reaction and determination. Energy activation and its

determination. Assumptions of activated complex theory. Fast reactions with examples. Polymers and their classification

12 hours

Unit-II

Crystallization kinetics. Basic concepts. Different methods Models for determining the crystallization kinetics for simple systems. In situ studies on crystallization kinetics.

12 hours

Unit-III

Thermodynamic equilibrium for simple to complex systems. Equilibrium reactions. Equilibrium and non-equilibrium phases. Estimation of solubility and stability. Dissociation constant. Activation coefficient. Speciation constant. Isoelectric points. Thermal stability and Chemical stability.

12 hours

Unit-IV

Thermodynamic modeling basic concept. Thermodynamic variables. Equation of state for predicting partial molal standard state properties of the system species. Helgeson equation of state. Calculation of excess properties of ions. Simple models. Pitzer model, Setschenow model. Bromley Zemaitis model. Helgeson Kirkham Flowers (HKF) model. Theoretical Phase diagrams with examples. Yield diagrams. Commercial softwares - CALPHAD, ChemApp, SimuSage, OLI Systems - FactSage Interface, etc.

12 hours

References

1. J. Thoma and B.O. Bouamama, Modelling and Simulation in Thermal and Chemical Engineering: A Bond Graph Approach, Springer (2000)
2. H. L. Lukas, Suzana G. Fries, Bo Sundman, Computational thermodynamics: the CALPHAD method, Cambridge University Press, UK (2007)
3. Jrgen Gmehling, Brbel Kolbe, Michael Kleiber, Jrgen Rarey, Chemical Thermodynamics: For Process Simulation, John Wiley & Sons (2012).
4. Advanced Course on Thermodynamic Models: Fundamentals & Computational Aspects, www.cere.dtu.dk, Department of Chemical and Biochemical Engineering - Slttofts Plads - Building 229 - DK2800 Kgs. Lyngby..

MSS-14: Basics of Engineering Drawing and Graphics 4 Credits (Soft Course)
Total: 48 hrs

Unit I : Principles of Graphics

Two dimensional geometrical construction Concept of section planes - Conic sections, involutes and cycloids - Representation of three dimensional objects - Principles of projections - standard codes of principles.

12 hours

Unit II: Orthographic Projections

Projections of points, straight line and planes - ' Auxiliary projections ' - Projection and sectioning of solids -Intersection of surfaces - Development of surfaces.

12 hours

Unit III: Pictorial Projections

Isometric projections - ' Perspectives ' - Free hand sketching. Conversion of pictorial views of simple machine parts into orthographic views, conversion of orthographic views of simple machine parts into isometric views.

12 hours

Unit IV: Computer Graphics

Hardware - Display technology - Software - Introduction to drafting software

12 hours

References

1. Narayanan, K.L., and Kannaiah, P., " Engineering Graphics ", Tata McGraw-Hill Publishers Co., Ltd., 1992.
2. William M. Neumann and Robert F.Sproul, " Principles of Computer Graphics ",McGraw Hill, 1989.
3. Warren J. Luzzadder and John M. Duff, " Fundamentals of Engineering Drawing ",Prentice-Hall of India Private Ltd., Eastern Economy Edition, 1995.
4. Natarajan K.V., " A Text Book of Engineering Drawing ", Private Publication, Madras, 1990.
5. Mathur, M.L. and Vaishwanar, R.S., " Engineering Drawing and Graphics ", Jain Brothers,New Delhi, 1993.

Note:

1. Further details on Basics of engineering drawing and graphics are illustrated below since this is a new course introduced to enhance the industrial acceptability of the M.Tech Program.

MSS-15: Ceramics Science and Technology 4 Credits (Soft core) Total: 48 hrs

Unit I

Ceramics: Definition, scope, bonding and structure, crystal structure, defects, polymorphism, ceramics as a class of material, variations within ceramics; defect structures; chronological developments, structure of silicates; polymorphic transformations, raw materials.

Unit II

Conventional Ceramics: b) Refractories : Classification of Refractories, Modern trends and developments, Basic raw materials, Elementary idea of manufacturing process technology, Flow diagram of steps necessary for manufacture, basic properties and areas of application. c) Whitewares : Classification and type of Whitewares, Elementary idea of manufacturing process technology including body preparation, basic properties and application areas. d) Ceramic Coatings : Types of glazes and enamels, Elementary ideas on compositions, Process of enameling & glazing and their properties. e) Glass : Definition of glass, Basic concepts of glass structure, Batch materials and minor ingredients and their functions, Elementary concept of glass manufacturing process, Different types of glasses.Application of glasses. f) Cement & Concrete : Concept of hydraulic materials, Basic raw materials, Manufacturing process, Basic compositions of OPC. Compound formation,setting and hardening.Tests of cement and concrete.

Unit-III

Elementary ideas about the raw materials used in pottery, Heavy clayweres, Refractoriers, Glass, Cement, Industries. Raw materials clays and their classification, Quartz, Polymorphism of quartz, Feldspar and its classification, Talc, Steatite and Mica. Fabrication methods: Packing of Powders, Classification and scope of various fabrication methods. Dry and semi dry pressing. extrusion, Jiggering & jollying, Slip casting HP & HIP. Drying & Firing of ceramics: Biscuit firing and glost firing, fast firing technology, action of heat on triaxial body, Elementary ideas of various furnaces used is ceramic industries.

12 hours

Unit-IV

Advanced ceramics: Bio-ceramics, Space ceramics, Automotive ceramics, Electronic ceramics, Superconducting ceramics, Elementary ideas of their preparation and applications.

12 hours

References

1. F.H Norton, Elements of Ceramics, Addison-Wesley Press (1974)
2. M.W. Barsoum, Fundamentals of Ceramics, McGraw-Hill (2003)
3. W.D Kingery, Introduction to Ceramics, Wiley & Sons (1976).
4. Lawrence H. Van Vlack, Physical Ceramics for Engineers, Addison-Wesley Publishing (1964).
5. F. Singer and S.J. Singer, Industrial Ceramics, Chapman & Hall, UK (1963)
6. Richerson D. W., 'Modern Ceramic Engineering - Properties Processing and Use in Design, 3rd Edition, CRC Press, 2006
7. Chiang Y.M., Birnie D. P., Kingery W.D., Physical Ceramics: Principles for Ceramic Science and Engineering, John Wiley, 1997
8. James E. Shelby., 'Introduction to Glass Science and Technology' 2nd Edition, The Royal Society of Chemistry Publications, 2005

MSS-16: Materials for Renewable Energy and Storage 4 Credits (Soft core)

Total: 48 hrs

Unit-I

Introduction to new generation of materials and nano-engineering of their structures for sustainable energy economy. Contribution to high performance renewable energy production, storage, conversion and usage. Solar grade glass; (a) properties-transparency, emissivity and reflectivity, (b) manufacturing- Flat glass for PV & CSP, tube glass for Evacuated Tube Receiver (ETR) and Collector (ETC).

12 hours

Unit-II

Solar Photo Voltaic (PV) cells: Single and multi-crystalline silicon solar cells, amorphous silicon, thin film; Cd-Te, CIGS, CZTS, nano-, micro-, poly-Si. Transparent conducting coating, Multi-junction, solar PV concentrator, flexible solar cells, Emerging PV; dye sensitized, other organic, and quantum dot cells. Nano-engineered materials.

12 hours

Unit-III

Materials for Concentrated Solar Power (CSP): Reflector materials; glass, metal, polymer and film. Receiver and collectors; absorptive coating and anti-reflective coating. Materials and shapes for thermal storage, Lithium ion Batteries. **12 hours**

Unit-IV

Fuel cells; materials and construction; PEM Fuel Cell(FC), AFC, PAFC, MCFC, SOFC. Catalysts for electro catalysis, fuel reformer and water splitting. **12 hours**

References

1. Vielstich, W., et al. (eds.) (2009). Handbook of fuel cells: advances in electrocatalysis, materials, diagnostics and durability. 6 vol. Hoboken: Wiley, 2009
2. Francis de Winter, Solar Collectors, Energy Storage, and Materials (Solar Heat Technologies), MIT Press, USA (1991)
3. David S. Ginley, David Cahen, Fundamentals of Materials for Energy and Environmental Sustainability, Cambridge University Press (2011)
4. Materials, Electronics, and Renewable Energy Part III Physics, Small lecture theatre, Cavendish Laboratory lecturers: David MacKay and Neil Greenham
5. Fuel Cell Handbook,
6. Introduction to Fuel Cell Technology

MSS-17: Basics of Nanotechnology

4 Credits (Soft core)

Total: 48 hrs

Unit-I

Introduction to nanotechnology basics, definition. History of nanotechnology. Nanotechnology in relation to other branches of science. Structure of solids crystalline and non-crystalline. Types of common materials and advanced materials inorganic, organic, biologic. Types of nanomaterials depending upon their properties electronic, semiconductors, superconductors, super ionic, magnetic, optic, opto-electronic, spintronics, lasers, photonics, ceramics, bio ceramics, biomedical, biosensors, bio imagers, photocatalysts, quantum dots. **12 hours**

Unit-II

Basic properties of materials and the instrumentation used to study these properties. Size effect of materials on properties. Quantization effect on the properties of materials with examples. Nanocomposites and their applications in modern technology. Nanotubes carbon nanotubes and other nanotubes. Nanomaterials natural and synthetic. Nanocomposites and Nanohybrid materials. **12 hours**

Unit-III

Nanomaterials synthesis techniques nanoengineering of materials. Bottom up and Top down routes. Solution, Melt and Gas Processing of nanomaterials. Nature inspired processes. **12 hours**

Unit-IV

Nanomaterials characterization X-rays, Spectroscopic - infrared, UV-Vis, Laser Raman, Photoluminescence, Electron Microscopic techniques, Thermal analysis, surface

characteristics, light scattering methods, gas adsorption, magnetic susceptibility, conductivity, band gap calculations. Nanotechnology in modern technology in relation to electronic, biological, consumer and domestic applications. Applied nanobiotechnology and nanobiomedical science drug delivery, drug targeting, biosensors, bioimaging, neutron capture therapy.

12 hours

References

1. Bharath Bhusan, Springer Handbook of Nanotechnology, 3rd edition, Springer-Verlag (2009)
2. CNR Rao and T. Cheetham, Chemistry of Nanomaterials : Synthesis, Properties and Applications, Wiley & Sons (2005)
3. Hari Singh Nalwa, Encyclopedia of Nanotechnology, American Scientific Publishers (2004)
4. K. Byrappa and M. Yoshimura, Handbook of Hydrothermal Technology, 2nd edition, Elsevier (2012)
5. K. Byrappa and T. Adschiri, Hydrothermal Technology for Nanotechnology, Progress in Crystal Growth and Characteriation of Materials, Volume 53 (2007) pp.117-166.
6. K. Byrappa and M. Yoshimura (Editors): Special Edition of Journal of Materials Science, Volume 41, No.6 (2006).
7. K. Byrappa and T. Adschiri (Editors), Special Edition of Journal of Materials Science, Volume 43, No.7 (2008).
8. Charles P. Poole Jr. and Franks J. Qwens, Introduction to Nanotechnology, Wiley & sons (2003)

MSS-18 Enterprise Architecture

4 Credits (Soft core)

Total: 48 hrs

Unit I:

1. What is Enterprise Architecture?
2. History of Architecture in the context of engineering discipline
3. History of Enterprise Architecture
4. Why Enterprise Architecture?
5. What are the various Enterprise Architecture Frameworks?
6. Why Zachman Framework an Ontology?

Unit II:

1. Growing enterprise opportunities
2. Enterprise Disorders & limitations of current approach
3. Zachman Framework & Primitive Models
4. Sample – Strategy Primitive model
5. Sample – Business process primitive models
6. Busting biggest myth about Enterprise Architecture
7. Summary

Unit III:

1. Case study1 Sample Problem–Application Rationalization arising out of Acquisition & Merger
2. Identify the cells responsible for the problem
3. Identify the primitives for each cell
4. Creating the composite models – 11 departments

5. Sample – Finance Department models, composites, relationship matrix, performance indicators
6. Composite Models & Traceability
7. Relationship Matrix, GAP Analysis & Application Maturity Index
8. Current status of Baseline Architecture v1.0
9. Timeline, Effort, Cost Estimations & Benefits

Unit IV:

1. Sample case 2 : Which cells are key to define the Disaster Recovery Plan
2. Sample case 3 - HR: Resource Request Process
3. Visual diagnosis for effective enterprise treatment
4. Use of Enterprise Architecture by Governments and Government Departments
5. Use of Enterprise Architecture in Telecom Industry
6. Use of Enterprise Architecture in Healthcare
7. Use of Enterprise Architecture in Banking & Finance
8. Use of Enterprise Architecture in Manufacturing

How to use Enterprise Architecture in managing Universities & Colleges

MSS 19: Chemical Engineering 4 credits (Soft core) Total: 48 hours

Unit I:

Introduction, concepts of Unit operation and Unit Processes, Units and Dimensions, Dimensional Analysis.

12 hours

Unit II:

Mass Transfer: Molecular and convective mass transfer, mass transfer equations. Heat Transfer: Heat transfer by conduction, convection and radiation, heat transfer equations, Heat exchangers. Momentum Transfer: Fluid flow, pipes and tubes, fittings, valves and pumps, measurement of pressure and flow.

12 hours

Unit III:

Size reduction, size separation, sedimentation, centrifugation, filtration, crystallization, Extraction, drying, evaporation, distillation. Sedimentation, Centrifugation, filtration, extraction, drying, evaporation, (Adsorption and partition) Techniques, Column thin layer and paper chromatography and their application. Criteria of purity-Melting and Boiling point. Surface tension, Viscosity.

12 hours

Unit IV:

Reaction Engineering: Reaction kinetics, order of reaction, Batch, fed batch and continuous reactors, reactor design. Process Control: Proportional, integral, derivative and PID controllers, stability of controllers. Basics of Flow sheet, material balance, materials economics, safety analysis

12 hours

1. Badger Walter, L. and Banchemo Julis, T., Introduction to Chemical Engineering, Tata McGraw Hill Pub. 1997.
2. McCabe, W.L., Smith, J.C. and Harriot, H.P., Unit Operations in Chemical Engineering, Mc Graw Hill, 2009
3. Jonson, C.D., Process Control and Instrumentation Technology, Prentic Hall India, 2003
4. Peters, Max S., and Timmerhaus, Klauss D., Plant Design and Economics for Chemical Engineers, Mc Graw Hill, 1991

MSS-20: Advanced X-ray Diffraction Studies 4 Credits (Hard core) Total: 48 hrs

Unit-I

Powder Diffraction Methods and calibration techniques: The modern Automated diffractometer: Applications of the Powder Method: Qualitative phase analysis: Crystallography and space group analysis: Indexing and lattice parameter determination, refinement and identification: Powder pattern calculation :Crystal structure determination - The Rietveld method.:

12 hours

Unit-II

Single Crystal Methods : Quantitative X-ray Diffraction: Interaction of X-rays with matter: absorption and EXAFS (time?): X-ray reflectometry analysis: Small Angle Scattering (5 lectures): Patterson Function: Pair correlation functions and linkage to structure function: Application to spherical, elliptical and needle shape inclusions":

12 hours

UNIT-III

Debyes function: Application to amorphous structures, nano-composites: (20 lectures) : Particle size and strain analysis line profile and Fourier techniques : Texture, Micro-texture and Residual stress "Pole figure in x-ray (single crystal and area detector).

12 hours

Unit-IV

Electron Diffraction (Orientation Imaging Microscopy): Fourier Analysis of Distributions: Euler angle definition of orientation space: Orientation Distribution Function: Fourier analysis of Orientation Distribution Function and quantification of texture: Stress (residual stress analysis)

12 hours

References

1. R. Jenkins and R. L. Snyder, Introduction to X-ray Analysis Diffractometry, John Wiley and Sons (1996):
2. Jens Als-Nielsen and Des Mc Morrow, Elements of Modern X-ray Physics, 2nd Edition Wiley & Sons (2011)
3. D. L. Bish and J. E. Post ed., Modern Powder, Diffraction Reviews in Mineralogy Vol. 20, Mineralogical Society of America, (1989).:
4. John Mc. Cowley Diffraction Physics, Klug and Alexander, X-ray Diffraction Procedures, J. Wiley and Son, New York (1972)
5. B. D. Cullity, Elements of X-ray Diffraction, Addison Wesley Publishing Company, Reading, Mass. (1956)

MSS -21: Analytical and Inorganic Chemistry 4 credits (soft core) Total: 48 hours

Unit I: Statistical Treatment of Analytical Data:

Limitations of analytical methods. Classification of errors:- Systematic errors- Sources, effects and their reduction. Random errors - Sources and distribution. Accuracy and precision. Fundamentals of chromatography: General description- Definition, terms and parameters used in chromatography (RF- value, retention volume and time). Ion-exchange chromatography (ICH), The synthesis of novel cation (Amberlite TR-1204) and anion exchange resin materials and its application in ion exchange chromatographic separation of the components from the reaction mixture. Potentiometry and conductometry: Theory, principle of working and few applications. Ionic bond and covalent bond: Properties of ionic substances, structure of ionic crystals, Hybridization, VSEPR concept to explaining the structure of simple molecules of materials. **12 hours**

Unit II: Physical Chemistry

Application of physical chemistry

Application of phase rule to two and three component systems. Concepts of entropy and free energy. Partial molar volume and its determination by density measurements. Symmetry elements and symmetry operations with examples of simple molecule materials. X-ray diffraction, Bragg equation and Miller indices. Order of a reaction and its determination. Energy of activation and its determination. Assumption of activated complex theory. Fast reactions with examples, polymers and their classification. Arrhenius theory of strong and weak electrolytes. Corrosion and its prevention. Law of photochemistry. Quantum yield and its determination. **12 hours**

Unit III:

Photodegradation of materials, Photocatalyst- ZnO, TiO₂, principle of photocatalyst, application of ZnO, TiO₂ in the photo degradation of various types, pesticides and in industrial effluents. Effect of photodegradation on chemical oxygen demand in drinking water and in industrial waste water. Photophysical properties of materials; Theory, instrumentation, and applications of fluorescence, characteristic of fluorescence, resonance fluorescence, sensitized fluorescence, quenching of fluorescence. Theory, principle, and applications of phosphorescence. **12 hours**

Unit IV: Organic Chemistry

Importance of natural products and synthetic products of organic origin materials in industry, pharmaceutical, petroleum refinery and agricultural fields. Uses of Dyes, polymers (plastics) soaps and detergents in industry, drugs and cosmetics in pharmaceutical industries, waxes, coal tar from petroleum industry and pesticides, (insecticides, pesticides, herbicides, fumigicide etc.) **12 hours**

References:

- 1) Quantitative analysis. R.A.DAY and A.L. Underwood 6th edition prentice hall, Inc 1999
- 2) Principle of Instrumental analysis, D.A. Skoog, F.J. Holler and T.A. Nieman, 5th edition. Thomson Asia pvt.Ltd. Singapore 1998
- 3) Analytical chemistry. G.D. Christian 5th edition 2001, John-wiley and sons Inc. India

- 4) Chemical kinetics. KJ. Laidler
- 5) Chemical kinetics. Moore and Pearson
- 6) Kinetics and Mechanism of Chemical Transformation by J. Rajaram and J.C. Kuriacose.
- 7) Advances in Photochemistry - Rohatgi Mukherjee
- 8) Principle and applications of Photochemistry – R.P. Wayne, Elsevier, New York, (1970).
- 9) Elements of physical chemistry – Glass stone and Lewis
- 10) Encyclopedia of chemical technology – Kirk-othmer series
- 11) Inorganic chemistry – JE Huheey
- 12) Chemical Kinetics –L.K. Jain.
- 13) Physical chemistry by PW. Atkins, ELBS, 4th edition, Oxford University press (1990).

MSS-22 Semiconductor optoelectronics 4 credits (Soft Core) Total 48 hours

Unit-I: Review of Semiconductor Device Physics:

Energy bands on solids, the E-k diagram, Density of states, Occupation probability, Fermi level and quasi Fermi levels, p-n junctions, Schottky junction and Ohmic contacts. Semiconductor optoelectronic materials, Bandgap modification, Heterostructures and Quantum Wells.

12 hours

Unit –II: Interaction of photons with electrons and holes in a semiconductor:

Rates of emission and absorption, Condition for amplification by stimulated emission, the laser amplifier.

Semiconductor Optical Amplifiers & Modulators: Semiconductor optical amplifiers (SOA), SOA characteristics and some applications, Quantum-confined Stark Effect and Electro-Absorption Modulators.

12 hours

Unit –III: Semiconductor Photon Sources:

Electroluminescence. The LED: Device structure, materials and characteristics. The Semiconductor Laser: Basic structure, theory and device characteristics; direct modulation. Quantum-well laser; DFB-, DBR- and vertical-cavity surface-emitting lasers (VCSEL); Laser diode arrays. Device packages and handling.

12 hours

Unit-IV: Semiconductor Photodetectors:

Types of photodetectors, Photoconductors, Single junction under illumination: photon and carrier-loss mechanisms, Noise in photodetection; Photodiodes, PIN diodes and APDs: structure, materials, characteristics, and device performance. Photo-transistors, solar cells, and CCDs. Optoelectronic integrated circuits – OEICs.

12 hours

Reference

1. B. E. A. Saleh and M. C. Teich, Fundamentals of Photonics, John Wiley & Sons, Inc., 2nd Ed. (2007), Ch.16, 17, and 18.
2. P. Bhattacharya, Semiconductor Optoelectronic Devices, Prentice Hall of India (1997).
3. J. Singh, Semiconductor Optoelectronics: Physics and Technology, McGraw-Hill Inc. (1995)
4. G. Keiser, Optical Fiber Communications, McGraw-Hill Inc., 3rd Ed. (2000), Ch.4,6.
5. A. Yariv and P. Yeh, Photonics: Optical Electronics in Modern Communications, Oxford University Press, New York (2007), 6th Ed.

Ch.15-17.

6.J. M. Senior, Optical Fiber Communication: Principles and Practice, Prentice Hall of India, 2nd Ed.(1994), Ch.6-8.

MSS-23 Analytical and Inorganic Chemistry 4 credits (Soft Core) Total: 48 hours

Unit – 1.1: Statistical treatment of Analytical data

Meaning and analytical prospective, scope and function of analytical chemistry. Analytical Problems and Their Solution. The nature, trends and limitation of analytical methods and procedure. Assessment of analytical data- definitions and basic concepts, The nature and origin of errors. Classification of errors- systematic errors- source, effects and their reduction. Random errors- source and distribution. Accuracy and precision. Measures of central tendency and variability. Reliability of results-confidence interval. Comparison of results- Student's t-test, comparing the two means and standard deviations-F-test, t-test and paired t-test. Rejection of a result-Q-test. Fundamentals of chromatography- Overview of analytical separations, The problem with simple separations, classifying analytical separations, general description- definition of chromatography, terms and parameters used in chromatography (RF- value, retention volume, retention time, Chromatographic Resolution). Ion - Exchange chromatography (ICH), the synthesis of novel cation (Amber TR - 1204) and anion exchange resin materials and its application in ion exchange chromatographic separation of the components from the reaction mixture. *Instrumental methods of analysis*: Theory, Instrumentation and applications of Colorimetry, Potentiometry , Conductometry, pH metry and Flame photometry. **12 hours**

Unit – 1.2: Basics of inorganic chemistry

Ionic bond and covalent bond: properties of ionic substances, structure of ionic crystals, salient feature of valence bond theory, Sigma and pi bonds shape of the simple inorganic compounds, Concepts of hybridization and to illustrate planar, tetrahedral square planer, and octahedral geometries, Molecular orbital theory and its application to form homo (H₂ N₂ & O₂) and hetero (HF,NO) diatomic molecules, VSEPR concept to explaining the structure of simple molecules of materials. Modern concepts of acids and bases: Arrhenius, Bronsted-Lowry, Lewis, Lux-Flood and solvents system concepts. Survey of essential and trace elements in biological system. **12 hours**

Unit – II: Physical Chemistry

Application of Physical Chemistry:

Statement of Gibb's phase rule and explanation of the terms involved, Thermodynamic derivation of phase rule, Phase diagram of one component system – water system, Condensed phase rule, Application to phase rule to two and three component systems- Eutectic Pb – Ag system and Fe –C system. Concepts of entropy and free energy. Partial molar volume and its determination by density measurements. Symmetry elements and symmetry operations with examples of simple molecule materials. X – Diffraction, Brag equation and miller indices. Order of a reaction and its determination. Energy of activation and its determination. Assumption of activated complex theory. Fast reactions with examples, polymer and their classification. Arrhenius theory of strong and weak electrolytes, Assumptions of Debye-Huckel theory of strong electrolytes. Electrochemical theory of corrosion, Galvanic series,

Types of corrosion- Differential metal corrosion, Differential aeration corrosion (Pitting and water line corrosion), Stress corrosion (caustic embrittlement in boilers), Factors affecting the rate of corrosion Corrosion control: Inorganic coatings – Anodizing and phosphating, Metal coatings – Galvanizing and Tinning, Corrosion inhibitors, cathodic protection. Laws of photo chemistry. Quantum yield and its determination. **12 hours**

Unit – III: photodegradation of materials, photocatalyst – ZnO, TiO₂:

Principle of photo catalyst, application of ZnO, TiO₂ in the photodegradation of various types, pesticides and its industrial effluents. Effects of photo degradation on chemical oxygen demand in drinking water and in industrial waste water. Photophysical properties of materials: theory, instrumentation, and applications of fluorescence, characteristics of fluorescence, resonance fluorescence, sensitized fluorescence, quenching of fluorescence. Theory, principle and applications of phosphorescence. **12 hours**

Unit – IV: Organic chemistry

Importance of natural products and synthetic products of organic origin materials in industry, pharmaceutical, petroleum refinery and agricultural fields. Use of dyes- Colour and constitution, classification, dyeing method and their industrial importance, polymers & plastics - Introduction, biodegradable and non-biodegradable polymers and their industrial importance, plastics (uses and effects on environment), natural and synthetic rubbers, polyamides and poly esters like nylon, decron, terelyne. Thermoplastics - poly carbonates, poly acrylates in lens applications, polyurethanes and conducting polymers. Soaps and detergents - Production and their cleansing action, drugs - Basic concepts, classification, sources, the requirement of an ideal drug and cosmetics - Detailed study of formulations and manufacturing of cream and lotions, lipstick and nail polish, shampoos, hair dyes and tooth pastes and pharmaceutical applications, waxes, coal tar from petroleum industry and pesticides, (insecticides, pesticides, herbicides, fumigicide etc.) **12 hours**

References:

1. *Quantitative analysis*. R. A. DAY and A.L. Underwood 6th edition, prentice hall, inc 1999.
2. *Principple of instrumental analysis*, D. A. Skoog, F. J. Holler and T.A. Nieman, 5th edition, Thomson Asia pvt .ltd, Singapore, 1998.
3. *Analytical Chemistry*, G.D. Christian, 5th edition, 2001, John-Wiley and sons, Inc. India
4. *chemical kinetics*, K.J. Laidler
5. *Chemical kinetics*, Moore and Pearson
6. *kinetics and Mechanism of chemical transformation* by J. Rajaram and J.C. Kuriacose.
7. *Advance in photo chemistry*- Rohatgi Mukherjee
8. *Principle and application of photochemistry*- R.P. Wayne, Elsevier, New York, (1970)
9. *elements of physical chemistry*- Glass stone and lewis.
10. *Encyclopedia of chemical technology*- Kirck- Othmer Series
11. *Inorganic chemistry*- J.E. Hucheey
12. *Chemical kinetics* – L.K. Jain
13. *Physical chemistry* by P.W. Atkins, ELBS, 4th edition, Oxford University Press (1990)

List of experiments for Practical classes

I –Semester

MSH-7: Characterisation Lab-1 **4 credits (Hard core)**

48 hours

Experiments:

1) Crystal Models-Crystal symmetry-Basic forms, 7 crystal systems, Holhedral forms, axial ratio, axial angles, crystal faces indexing.

Important hemihedral and hemimorphic forms.

2) X-ray powder pattern method

3) UV-Visible spectrometer recording of a material and interpretation

4) (a) Microscopic identification of cells in fixed slides: plant cells, bacterial cells, fungal hypha and animal tissues. (b) Observing stages

of cell cycle – onion root ti (c) Microbial cell culture techniques – bacterial and fungal

(d) Gram staining

5) (a) Spread plate technique

(b) Streaking method

(c) Serial dilution of bacterial cultures (Ex:- isolation of bacteria from mud)

(d) Size determination of biomolecules using microscopic techniques.

6) Particle size determination by optical diffraction method.

7) Particle size determination by X-ray diffraction technique.

8) Thermistor-temperature variation.

II –Semester

MSH-8: Characterisation Lab-2 **4 credits (Hard core)**

48 hours

Experiments:

1) FTIR and ¹H NMR recording of a materials/organic/inorganic compounds and interpretation

2) Determination of the birefringence of the given sample using the Abbe

Refractometer/Prism coupler.

3) Measurement of the band gap energy of the given semiconductor by the method of four probes

4) Determine the refractive index and thickness of polymer films using Metricon prism coupler

5) Free volume estimation of a polymer by positron annihilation lifetime spectroscopy

6) Measure the AC impedance of the given composite and construct the Bode-Bode plot

7) Well diffusion method ,Disk diffusion method, Antimicrobial activity of nanoparticles using.

8) Isolation of plasmid DNA from *E.coli*

9) Isolation of genomic DNA from *E.coli*

10) Transformation of cell – heat shock technique

11) Isolation of DNA and spectrophotometric analysis

III –Semester

MSH-9: Characterisation Lab-3

4 credits (*Hard core*)

48 hours

Experiments:

1. Paramagnetic susceptibility of the given salts.
2. Hall effect in metals/semi conduction.
3. Crystallization
4. Distillation.
5. Fractional crystallization.
6. Synthesis of aspirin, antipyrine
7. Synthesis of liquid crystals.
8. Thin layer chromatography and column chromatography
9. Isolation of DNA from animal tissues
10. Isolation of DNA from plant tissues
11. Isolation of mRNA from animal tissues
12. Restriction enzyme digestion
13. DNA ligation
14. Agarose gel electrophoresis (AGE)
15. SDS PAGE electrophoresis
16. Biological materials characterization by X-ray diffraction studies.

IV- Semester

MSH-10 Characterisation Lab.-4

4 credits (*Hard core*)

48 hours

Experiments (for B.Sc., students)

1. Preparation of simple materials: Crystals, glasses, polymers, and composites by hydrothermal, solvothermal, Flux, gel, melt and sol-gel techniques.
2. Synthesis of Au, Ag, Cu NPs by Chemical/Biological methods.
3. Preparation of materials exposed to microwave radiations.
4. Preparation of pellets of materials.
5. Fermi energy estimation of conductors/Semi conductors using Bridge.
6. Morphological studies of prepared materials.
7. Morphological changes in materials due to irradiation.
8. Animal cell culture techniques – primary cell culture, sub-culturing.
9. Size determination of bio-materials using microscope and Zeta potential.
10. Protein adsorption in bio-materials.
11. Cell bio-material interpretation.

V-Semester

MSH-11 Characterisation Lab.-5

4 credits (*Hard core*)

48 hours

Characterization Lab-5 is For B.Sc, Students

- 1) Preparation of carbon nanoparticles and X-ray investigation
- 2) FTIR and UV-Visible studies of polymers (both natural and man-made)
- 3) Preparation of polymer film and x-ray investigation
- 4) X-ray studies of natural polymers like silk, cotton, hemp and jute
- 5) AFM Study of a polymer film/composite material/Nanomaterials
- 6) SEM study of a polymer film/composite material/Nanomaterials
- 7) Conductivity of a polymer film

PROJECT WORK

MSH-12: Minor project during III Semester for non B.Sc. students and in V semester for B.Sc. students. : 6 credits

MSH-13 Major project during IV semester for non-B.Sc. students and in VI Semester for B.Sc. students. : 10 credits
